

(M) – , 1

n, / m / .

Mr (Ar).

$r(\text{CH}_4) = \text{Ar}(\text{C}) + 4 \text{Ar}(\text{H}) = 12 + 4 = 16$
 $M(\text{CH}_4) = 16 / \dots 16 \text{ CH}_4 \quad 6,02 \cdot 10^{23}$

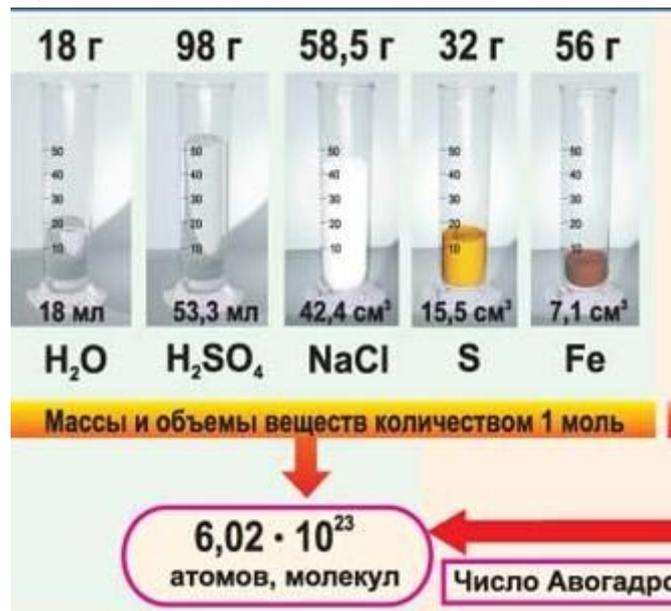
() n, : $M = m/n$

: $n = m/M$

: $m = n \cdot M$

, ... Mr Ar.

m.





CH₄

2H₆,

2

:

$$M(\text{CH}_4) = 16 \text{ / } ;$$

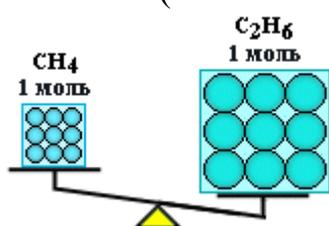
$$M(\text{C}_2\text{H}_6) = 2 \cdot 12 + 6 = 30 \text{ / } .$$

:

$$m(\text{CH}_4) = 2 \cdot 16 \text{ / } = 32 \text{ ;}$$

$$m(\text{C}_2\text{H}_6) = 2 \cdot 30 \text{ / } = 60 \text{ .}$$

, - , , ...
()



$$n(\text{CH}_4) = n(\text{C}_2\text{H}_6), \quad m(\text{CH}_4) < m(\text{C}_2\text{H}_6)$$

n

<p>1. () ,</p> <p>0,5 ?</p> <p>: $n(\text{Fe}) = 0,5$</p> <p>: $m(\text{Fe}) - ?$</p> <p>:</p> <p>$m = M \cdot n$</p> <p>$M(\text{Fe}) = Ar(\text{Fe}) = 56 \text{ / } ()$</p> <p>$m(\text{Fe}) = 56 \cdot 0,5 = 28$</p> <p>: $m(\text{Fe}) = 28$</p>
<p>2. () $12,04 \cdot 10^{23}$ Ca ?</p> <p>: $N(\text{CaO}) = 12,04 \cdot 10^{23}$</p> <p>: $m(\text{CaO}) - ?$</p> <p>:</p> <p>$m = M \cdot n, n = N/Na,$</p> <p>,</p> <p>$m = M \cdot (N/Na)$</p> <p>$M(\text{CaO}) = Ar(\text{Ca}) + Ar(\text{O}) = 40 + 16 = 56 \text{ / }$</p> <p>$m = 56 \cdot (12,04 \cdot 10^{23} / 6,02 \cdot 10^{23}) = 112$</p> <p>: $m = 112$</p>